Additions and Corrections

Electronic Structure and Properties of Transition Metal−Benzene Complexes [*J. Am. Chem. Soc.* **2001**, *123*, 3799−3808]. RAVINDRA PANDEY, BIJAN K. RAO,* PURUSOTTAM JENA, AND MIGUEL ALVAREZ BLANCO

We have found an inconsistency in our recent paper in the definition of the electron affinity of the metal atom, metal−benzene, and metal−$(benzene)_2$ complexes in Figure 8. A modified Figure 8 is presented here. The definition of the electron affinity to be used is given by E.A. = *E*(anion) − *E*(neutral). With this definition, a negative electron affinity means that the anion is stable. Consequently, the variations of the electron affinities of the metal−benzene complexes and the metal−$(benzene)_2$ complexes across the 3d series (Sc−Ni) are similar in nature and both Ni−benzene and Ni−$(benzene)_2$ anions are unstable against autodetachment of the electron. The calculated electron affinity of V−$(benzene)_2$ is +0.29 eV. The experimental study suggests that the V−$(benzene)_2$ anion may be unstable in the gas phase.

The remaining discussions in section III.F remain unchanged.

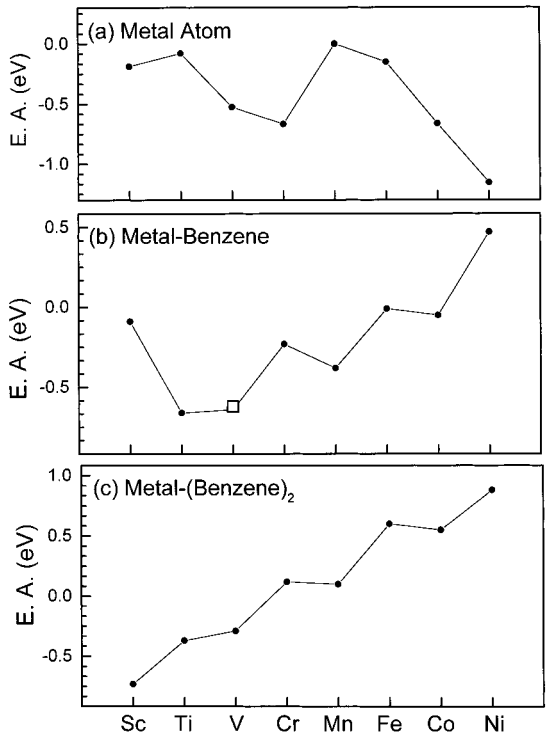

Figure 8.

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